

Acid and Base Nomenclature

- Acids with Monatomic Anion
 - Example: HCl
 - Start with the prefix “hydro”

 - Add the stem of the anion

 - Add the suffix “ic”

 - Always end with acid
Final Name: _____
- Acids with Polyatomic Anions that End in “-ate” (example: HClO₃)
 - Example: HClO₃
 - Add the stem of the anion

 - Add the suffix “ic”

 - Always end with acid
Final Name: _____
- Acids with Polyatomic Anions that End in “-ite”
 - Example: HClO₂
 - Add the stem of the anion

 - Add the suffix “ous”

 - Always end with acid
Final Name: _____

- We name acids differently than normal molecular substances
- How we name acids depends on the anion that is found within them.
 - Acids with Monatomic Anion (example: HCl)
 - Start with the prefix “_____”
 - Add the _____ of the anion
 - Add the suffix “_____”
 - Always end with acid
 - Acids with Polyatomic Anions that End in “-ate” (example: HClO₃)
 - Start with the _____ of the anion
 - Add the suffix “_____”
 - Always end with acid
 - Acids with Polyatomic Anions that End in “-ite” (example: HClO₂)
 - Start with the _____ of the anion
 - Add the suffix “_____”
 - Always end with acid
 - Don’t forget the helpful mnemonic device: “I _____ something _____ in the cafeteria.”
- Bases are named following the standard naming conventions for ionic compounds. There’s no need to learn anything new or special for them!