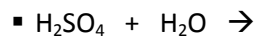
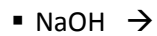
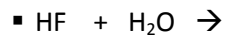
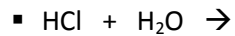


What Are Acids and Bases?

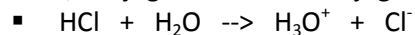
▪ Arrhenius Acids and Bases

- Write dissociation reactions for the following Arrhenius acids and bases:



▪ Bronsted-Lowry Acids and Bases

- For each of the following reactions, label the Bronsted-Lowry acid, base, conjugate acid and conjugate base.

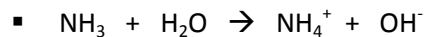


- Acid: _____

- Base: _____

- Conjugate Acid: _____

- Conjugate Base: _____



- Acid: _____

- Base: _____

- Conjugate Acid: _____

- Conjugate Base: _____

- Describe the following as monoprotic or polyprotic:

- HCl _____

- H_2SO_4 _____

- HBr _____

- H_3PO_4 _____

- Acids and bases are special types of electrolytes.

- Electrolytes are substances that _____ in water

- We have two different definitions of acids and bases:

- Arrhenius Acids and Bases

- _____ dissociate in water to produce an excess of _____

- The dissociation reaction can be written following the form:

- _____

- _____ dissociate in water to produce an excess of _____

- _____

- Bronsted-Lowry Acids and Bases

- _____ acids are substances that act as proton _____

- _____ acid is the substance that is produced after the Bronsted-Lowry accepts a proton.

- _____ bases are substances that act as proton _____

- _____ base is the substance that is produced after the Bronsted-Lowry acid donates a proton.

- Substances that can act as both an acid and a base are called _____.

- _____ acids – acids that contain and dissociate only one H^+

- _____ acids – acids that contain and dissociate multiple H^+