Concentration of Solutions Worksheet

Calculate the concentrations of the following solutions using both percent by mass and molarity. (Assume that the density of water is 1.00 g/mol)

1) 2.3 moles of sodium chloride in 0.45 liters of solution.
   a) Percent by Mass Concentration    b) Molarity

2) 1.2 moles of calcium carbonate in 1.22 liters of solution.
   a) Percent by Mass Concentration    b) Molarity

3) 98 grams of sodium hydroxide in 2.2 liters of solution.
   a) Percent by Mass Concentration    b) Molarity

4) 1.2 grams of hydrochloric acid in 25 mL of solution.
   a) Percent by Mass Concentration    b) Molarity

**Explain how you would make** the following solutions. (Follow the example on the website!)

5) 1.5 L of 2.00 M NaOH\(_{(aq)}\)

6) 0.75 L of 0.25 M Na\(_2\)SO\(_4\)\(_{(aq)}\)